

ought to be disregarded, but they do not forward any criterion that might be useful for recognizing, i.e. identifying, such data. Hence, the new method owes its simplicity to an incorrect elimination of the effect of side diffusion.

Moreover, problems arising due to the necessity of controlling the temperature of a sizeable bulk of liquid should be considered; especially near the surface the layer being measured may be cooled by evaporation to the overlying gas which is not saturated by the liquid vapour. With large volumes of liquid it also becomes more difficult to prevent mechanically induced convective fluxes and their fluctuations. Hence, the new method seems to be rather vulnerable because of serious doubts about its usefulness, in spite of the fact that it has already been popularized in at least four publications (Ju and Ho, 1985, 1986; Ho *et al.*, 1986; Ju *et al.*, 1986). According to Ju and Ho (1986) "the literature appears to be scattered with information which is confusing and often mutually contradictory. . . It appears that the controversy in large might be related to the problems associated with different experimental techniques." But this criticism fits their own work to perfection.

VÁCLAV LINEK  
VÁCLAV VACEK

Chemical Engineering Department  
Institute of Chemical Technology  
Suchbátarova 1905  
16628 Prague 6  
Czechoslovakia

## REFERENCES

- Akita, K., 1981, Diffusivities of gases in aqueous electrolyte solutions. *Ind. Engng Chem. Fundam.* 20, 89-94.  
Goldstick, T. K. and Fatt, I., 1970, Diffusion of oxygen in solutions of blood proteins. *Chem. Engng Prog. Symp. Ser.* 66, 101-113.  
Ho, C. S., Ju, L. and Ho, C. T., 1986, Measuring oxygen diffusion coefficients with polarographic oxygen electrodes II. Fermentation media. *Biotechnol. Bioengng* 28, 1086-1092.  
Ju, L. and Ho, C. S., 1985, Measuring oxygen diffusion coefficients with polarographic oxygen electrodes I. Electrolyte solutions. *Biotechnol. Bioengng* 17, 1495-1499.  
Ju, L. and Ho, C. S., 1986, The measurement of oxygen diffusion coefficients in polymeric solutions. *Chem. Engng Sci.* 41, 579-589.  
Ju, L., Livio, E. and Ho, C. S., 1986, Measuring oxygen diffusion coefficients in electrolyte solutions with polarographic oxygen electrodes. *Biotechnol. Bioengng Symp.* 15, 347-361.  
Linek, V. and Vacek, V., 1981, Chemical engineering use of catalyzed sulfite oxidation kinetics for the determination of mass transfer characteristics of gas-liquid contactors. *Chem. Engng Sci.* 36, 1747-1768.  
Nakanoh, M. and Yoshida, F., 1980, Gas absorption by Newtonian and non-Newtonian liquids in a bubble column. *Ind. Engng Chem. Process Des. Dev.* 19, 190-195.  
Yagi, H. and Yoshida, F., 1975, Gas absorption by Newtonian and non-Newtonian fluids in sparged agitated vessels. *Ind. Engng Chem. Process Des. Dev.* 14, 488-495.

Chemical Engineering Science, Vol. 42, No. 10, pp. 2488-2489, 1987.  
Printed in Great Britain.

0009-2509/87 \$3.00+0.00  
Pergamon Journals Ltd.

## Comments on improvements on a replacement for the logarithmic mean

(Received 13 February 1987; accepted 5 May 1987)

Dear sirs,

In a recent publication Paterson [1] gave an approximation to the logarithmic mean temperature difference (LMTD). In the range of interest for heat exchanger design, it gave close approximations to the LMTD and it also has the advantage that it approaches the correct limit as the two temperature differences approach one another. Furthermore, it is mathematically tractable and provides an explicit solution when the outlet temperature and the flow rate of one stream, say the cooling stream, are the unknowns.

Paterson's approximation  $\theta_{PM}$  to the LMTD is given as

$$\theta_{LM} \approx \theta_{PM} = \frac{2}{3} \theta_{GM} + \frac{1}{3} \theta_{AM} \quad (1)$$

where

$$\theta_{GM} = (\theta_1 \theta_2)^{1/2} \quad (2)$$

$$\theta_{AM} = \frac{1}{2}(\theta_1 + \theta_2). \quad (3)$$

Equation (1) may be substituted into the rating equation of a heat exchanger

$$Q = UA\theta_{LM} \quad (4)$$

which for given  $(Q/UA)$  may be simplified to result in a

quadratic involving  $\theta_1$  and  $\theta_2$  and, after some rather laborious manipulation,  $\theta_1$  is obtained as an explicit solution of  $\theta_2$

$$\theta_1 = \left( 7\theta_2 + 6 \frac{Q}{UA} \right) - 4 \sqrt{3 \left( \theta_2^2 + 2\theta_2 \frac{Q}{UA} \right)^{1/2}} \quad (5)$$

In the counter-current heat exchanger example shown by Paterson [1], the exit temperature may be calculated from  $\theta_1$ , as all the terms on the RHS of eq. (5) are known.

It is interesting to note that the mean temperature difference approximation given by Paterson [1] may be viewed as a weighted arithmetic mean of  $\theta_{GM}$  and  $\theta_{AM}$ . One could also derive an approximation based on the weighted geometric mean of  $\theta_{GM}$  and  $\theta_{AM}$

$$\theta_{CMI} = \theta_{GM}^a \theta_{AM}^b \quad (6)$$

where  $a + b = 1.0$ , and it would simplify the subsequent analysis if  $a = 2b$ .

Thus, one could try

$$\theta_{CMI} = \theta_{GM}^{2/3} \theta_{AM}^{1/3} \quad (7)$$

which will result in, after simplification,

$$\theta_{CM1} = \left( \frac{\theta_1 \theta_2^2 + \theta_1^2 \theta_2}{2} \right)^{1/3} \quad (8)$$

If eq. (8) is substituted into eq. (4),  $\theta_1$  may again be obtained as an explicit function of  $\theta_2$ :

$$\theta_1 = -\frac{\theta_2}{2} + \sqrt{\left[ \frac{\theta_2^2}{4} + \frac{2}{\theta_2} \left( \frac{Q}{UA} \right)^3 \right]} \quad (9)$$

The derivation and the form of eq. (9) are much simpler than for the case of eq. (5) given by Paterson [1].

The two approximations given above are based on the consideration of some weighted mean values of  $\theta_{GM}$  and  $\theta_{AM}$ .

A search in the literature revealed that the celebrated British chemical engineer, A. J. V. Underwood [2, 3], had also given an approximation to  $\theta_{LM}$ , in 1933, based directly on  $\theta_1$  and  $\theta_2$ :

$$\theta_{UM}^{1/3} = \frac{1}{2} (\theta_1^{1/3} + \theta_2^{1/3}) \quad (10)$$

While Underwood [3] probably derived eq. (10) with the aid of a slide rule, it is possible to fine-tune the exponential index to obtain a better match with  $\theta_{LM}$ . This results in

$$\theta_{CM2}^{0.3275} = \frac{1}{2} (\theta_1^{0.3275} + \theta_2^{0.3275}) \quad (11)$$

When either eq. (10) or (11) is applied to eq. (4) and solving for  $\theta_1$  in terms of  $\theta_2$  the following result is simply obtained:

$$\theta_1 = \left[ 2 \left( \frac{Q}{UA} \right)^n - \theta_2^n \right]^{1/n} \quad (12)$$

where  $n = \frac{1}{3}$  or 0.3275 depending on whether eq. (10) or (11) was used.

To compare the accuracies of the predictions of Paterson's approximation and those given in this work,  $\theta_{LM}$ ,  $\theta_{PM}$ ,  $\theta_{CM1}$ ,  $\theta_{UM}$  and  $\theta_{CM2}$  are evaluated and given in Table 1 for the same  $\theta_1$  and  $\theta_2$  values used by Paterson [1].

With reference to Table 1, it is observed that while eq. (8) provides a simpler solution than eq. (1), its accuracy is slightly inferior. However, the approximation given by Underwood in 1933 which results in a very simple solution as evidenced by eq. (12) also gives superior approximations to  $\theta_{LM}$  compared to that proposed by Paterson [1]. Furthermore, by slightly adjusting the exponential index to give  $\theta_{CM2}$  as modified in this work, the approximation gives an essentially exact match with  $\theta_{LM}$  in the range of  $\theta_1$  and  $\theta_2$  values considered. The application of eq. (12) to the worked example given by Paterson [1] is straightforward and need not be given here.

#### CONCLUSION

A slight modification of the approximation to  $\theta_{LM}$  given by Paterson [1] results in a simpler solution although its

Table 1. Comparison of the various approximations with  $\theta_m$ , the logarithmic mean

	10	10	10	10
$\theta_1$	10	10	10	10
$\theta_2$	15	20	50	100
$\theta_{LM}$	12.33	14.43	24.85	39.09
$\theta_{PM}$ [eq. (1)]	12.33	14.42	24.91	39.42
$\theta_{CM1}$ [eq. (8)]	12.33	14.42	24.66	38.03
$\theta_{UM}$ [eq. (10)]	12.33	14.42	24.88	39.24
$\theta_{CM2}$ [eq. (11)]	12.33	14.42	24.84	39.09

accuracy is slightly inferior. An approximation given by Underwood in 1933 [2, 3] was shown to give not only a very simple solution but also superior results. Furthermore, a slight modification to the Underwood approximation results in eq. (11) giving  $\theta_{CM2}$  which is in almost exact agreement with  $\theta_{LM}$ , the LMTD, in the range of  $\theta_1$  and  $\theta_2$  of interest.

The advantage of having available an approximation to the LMTD has already been pointed out by Paterson [1] and it is suggested that in situations where an approximation is required, eq. (11) should be seriously considered.

J. J. J. CHEN

Department of Chemical and  
Materials Engineering  
University of Auckland  
Private Bag, Auckland  
New Zealand

#### NOTATION

$A$	exchange area
$Q$	heat duty
$U$	overall heat transfer coefficient
$\theta$	temperature difference

#### Subscripts

1, 2	ends of the exchanger
AM	arithmetic mean
CM1	approximation given in this work, eq. (7)
CM2	approximation given in this work, eq. (11)
GM	geometric mean
LM	logarithmic mean
PM	approximation of given by Paterson [1], eq. (1).
UM	approximation given by Underwood [2, 3], eq. (10)

#### REFERENCES

- [1] Paterson, W. R., *Chem. Engng Sci.* **39**, 1635-1636 (1984).
- [2] Underwood, A. J. V., *Ind. Chemist.* pp. 167-170 (May 1933).
- [3] Underwood, A. J. V., *Chemical Engineering*, p. 192 (15 June 1970).